

DOUBLE DISPLACEMENT REACTIONS

Use **only COEFFICIENTS** to balance the following equations. All subscripts have already been added.

1. $\text{NaBr} + \text{AgNO}_3 \longrightarrow \text{NaNO}_3 + \text{AgBr}$
2. $\text{Ca}(\text{NO}_3)_2 + \text{K}_2\text{SO}_4 \longrightarrow \text{CaSO}_4 + \text{KNO}_3$
3. $\text{KCl} + \text{Pb}(\text{NO}_3)_2 \longrightarrow \text{KNO}_3 + \text{PbCl}_2$
4. $\text{HCl} + \text{FeO} \longrightarrow \text{FeCl}_2 + \text{H}_2\text{O}$
5. $\text{NaOH} + \text{H}_2\text{SO}_4 \longrightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$
6. $\text{H}_2\text{S} + \text{HgCl}_2 \longrightarrow \text{HCl} + \text{HgS}$
7. $\text{Ca}(\text{OH})_2 + \text{Fe}(\text{NO}_3)_3 \longrightarrow \text{Ca}(\text{NO}_3)_2 + \text{Fe}(\text{OH})_3$
8. $\text{Pb}(\text{NO}_3)_2 + \text{H}_2\text{SO}_4 \longrightarrow \text{PbSO}_4 + \text{HNO}_3$
9. $\text{Mg}(\text{OH})_2 + \text{HNO}_3 \longrightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2\text{O}$
10. $\text{Al}_2(\text{SO}_4)_3 + \text{NaOH} \longrightarrow \text{Al}(\text{OH})_3 + \text{Na}_2\text{SO}_4$